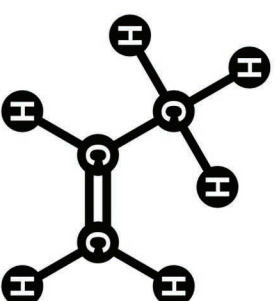
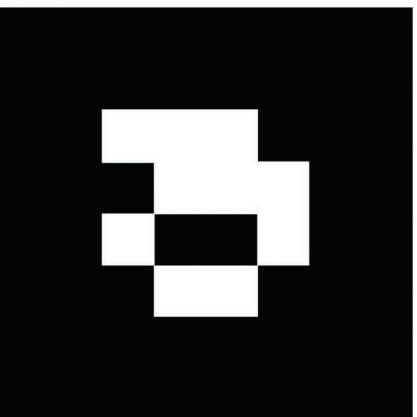


## $C_2H_4$ Ethylene

Imagine ethylene,  $C_2H_4$ , in three-dimensions. What are the hybridizations of the carbon atoms and what are the bond angles?

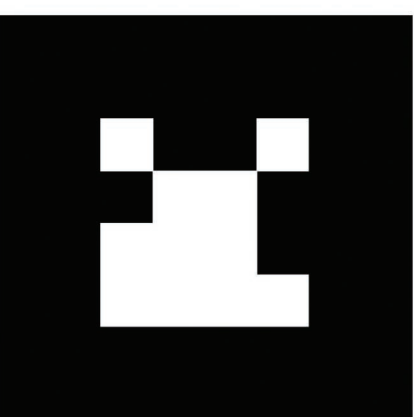
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## $CH_3CHCH_2$ Propylene

Imagine propylene,  $CH_3CHCH_2$ , in three-dimensions. What are the hybridizations of the carbon atoms and what are the bond angles?

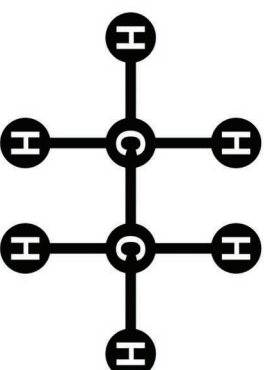
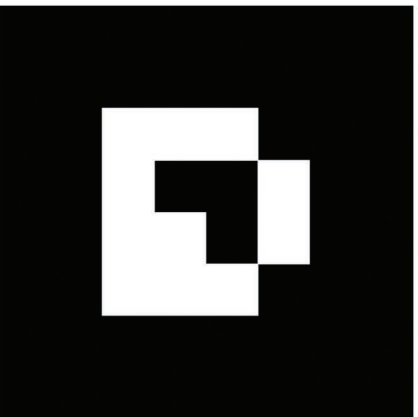
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## $C_2H_2$ Acetylene

Imagine acetylene,  $C_2H_2$ , in three-dimensions. What are the hybridizations of the carbon atoms and what are the bond angles?

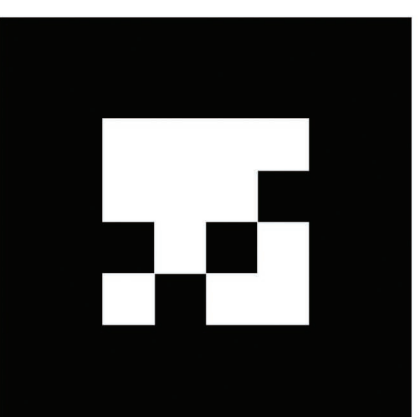
[www.SponholtzProductions.com](http://www.SponholtzProductions.com)

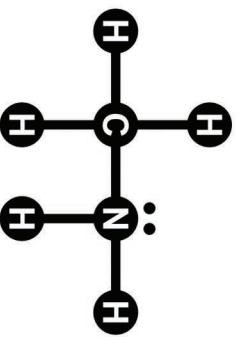


## $C_2H_6$ Ethane

Imagine ethane,  $C_2H_6$ , in three-dimensions. What are the hybridizations of the carbon atoms and what are the bond angles?

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### $\text{CH}_3\text{NH}_2$ Methylamine

Imagine methylamine,  $\text{CH}_3\text{NH}_2$ , in three-dimensions. What is the hybridization of the carbon and nitrogen atoms and what are the bond angles?

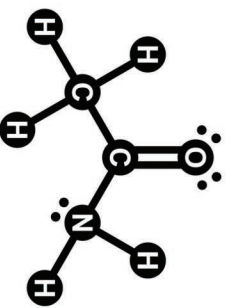
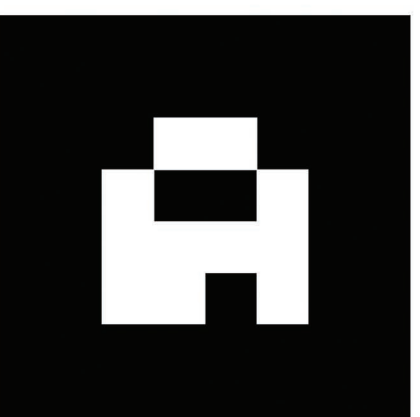
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### $\text{CO}_2$ Carbon Dioxide

Imagine carbon dioxide,  $\text{CO}_2$ , in three-dimensions. What are the hybridizations of the carbon and oxygen atoms and what are the bond angles? Are the two adjacent double bonds (pi bonds) perpendicular?

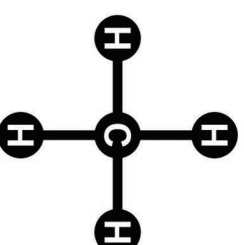
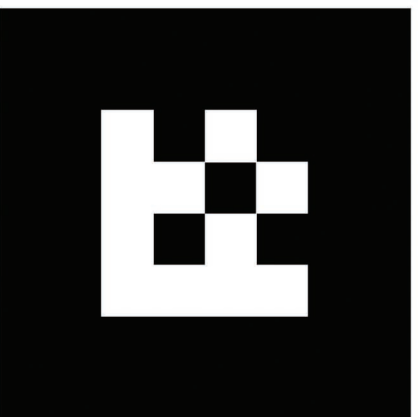
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### $\text{CH}_3\text{CONH}_2$ Acetamide

Imagine acetamide,  $\text{CH}_3\text{CONH}_2$ , in three-dimensions. What are the hybridizations of the carbon and nitrogen atoms and what are the bond angles?

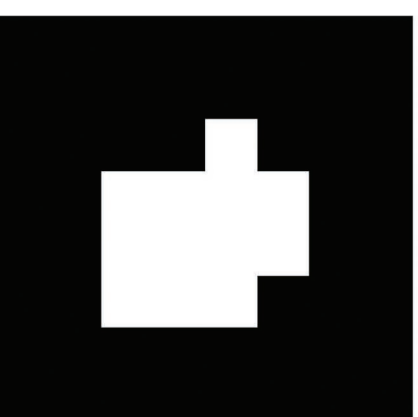
[www.SponholtzProductions.com](http://www.SponholtzProductions.com)

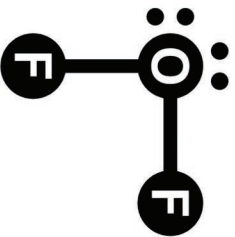


### $\text{CH}_4$ Methane

Imagine methane,  $\text{CH}_4$ , in three-dimensions. What is the hybridization of the carbon atom and what are the bond angles?

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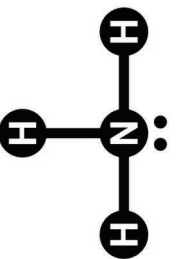
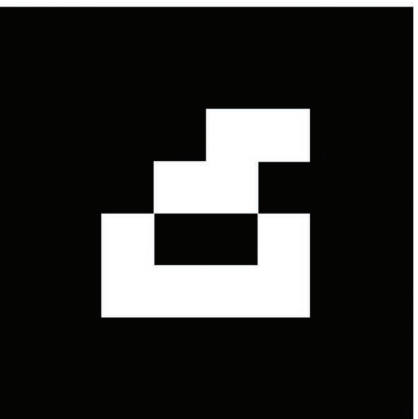




## OF<sub>2</sub> Oxygen Difluoride

Imagine oxygen difluoride, OF<sub>2</sub>, in three-dimensions and indicate any deviations from ideal bond angles.

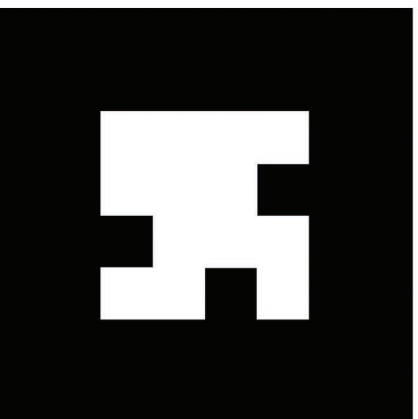
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## NH<sub>3</sub> Ammonia

Imagine ammonia, NH<sub>3</sub>, in three-dimensions. What is the hybridization of the nitrogen atom and what are the bond angles?

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## sp<sup>3</sup>, sp<sup>2</sup>, and sp Hybridized Carbon

Imagine a sp<sup>3</sup>, sp<sup>2</sup>, and a sp hybridized carbon atom in 3-D. How many hybridized and unhybridized lobes does each have?

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